

Herbert S. Harned and Calvin Calmon. The Properties of Electrolytes in Mixtures of Water and Organic Solvents. I. Hydrochloric Acid in Ethanol- and Isopropanol-Water Mixtures of High Dielectric Constant.

Page 1492. The sign on the right of equation (4) should be changed to +, that on the right of equation (5) to -. This error affects the results in columns headed E_c^0 and E_N^0 in the last two columns of Table III (page 1493).—HERBERT S. HARNED.

Frank T. Gucker, Jr., and Charles E. Moser. The Coefficient of Expansibility of Aqueous Solutions of Urea at 27.5° Calculated from the Densities at 25 and 30°.

Page 1558. Column 2. In equation 6 the last term should read:

$$\frac{1}{d_1} \left[\frac{\partial A_1}{\partial T} - 2A_1\alpha_1 + \frac{A_1}{d_1} \left(\frac{\partial A_1}{\partial T} \right) \right] c^3 + \dots \quad (6)$$

—FRANK T. GUCKER, JR.

M. S. Kharasch, Cheves Walling and Frank R. Mayo. The Addition of Hydrogen Halides to *cis*- and *trans*-2-Pentene.

Page 1562. In Table II the various temperatures marked 75, 78 and 80 indicate the temperature of an acetone-solid carbon dioxide bath and should be prefixed by a minus sign.—CHEVES WALLING.

Louis F. Fieser and William S. Johnson. Syntheses in the 1,2-Benzanthracene and Chrysene Series.

Page 1652. Column 1, line 5, for "5-Methyl-8-keto-3,4,5,6,7,8-hexahydrophenanthrene (IX)" read "5-Methyl-8-keto-3,4,5,6,7,8-hexahydro-1,2-benzanthracene (IX)."—L. F. FIESER.

Cheves Walling, M. S. Kharasch and F. R. Mayo. The Peroxide Effect in the Addition of Reagents to Unsaturated Compounds. XX. The Addition of Hydrogen Bromide to 2-Butyne and 2-Bromo-2-butene.

Page 1711. In footnote (4) for "see ref. 7" read "see ref. 9."—CHEVES WALLING.

Philip G. Stevens. The Rearrangement of α -Hydroxy Carbonyl Compounds.

Page 1714. Column 1, line 13, for "Fischer's discovery" read "the discovery by Fischer, Taube and Baer." Column 1, footnote (3), for "Fischer, *ibid.*" read "Fischer, Taube and Baer, *ibid.*"—PHILIP G. STEVENS.

James J. Lingane. Thermodynamic Significance of Polarographic Half-Wave Potentials of Simple Metal Ions at the Dropping Mercury Electrode.

Page 2102. "Due to an error in respect to concentration units the calculated average value of k , 0.005 mole fraction per microamp., given in the concluding paragraph is incorrect. The correct average value of k is 5×10^{-6} mole fraction per microamp., or 0.3 millimole per liter of amalgam per microamp. As a consequence the statements

in the concluding paragraph are incorrect. The rate of diffusion of the deposited metal atoms in the mercury drops is about equal to that of the reducible metal ions in the solution, and not much slower, as was previously concluded."—JAMES J. LINGANE.

M. L. Wolfrom, D. R. Myers and E. N. Lassetre. The Molecular Size of Starch by the Mercaptalation Method.

Page 2175. Column 2, line 1, in the formula for "%S + " read "%S \times ."—M. L. WOLFROM.

James J. Pyle, Leo Brickman and Harold Hibbert. Studies on Lignin and Related Compounds. XLIV. The Ethanolysis of Maple Wood; Separation and Identification of the Water-Soluble Aldehyde Constituents.

Page 2201. In column 1, after the formulas, insert: "Possible Role of Vanilloyl- and Syringoyl aldehydes in the Synthesis of Plant Pigments.—It was shown some years ago by Bülow and Sicherer⁹ that benzoyl aldehyde, in the presence of acids, readily undergoes condensation with a variety of phenols (resorcinol, phloroglucinol, etc.)."

(9) Bülow and Sicherer, *Ber.*, **34**, 3889 (1901).

—HAROLD HIBBERT.

Kinney Hancock and H. L. Lochte. Acidic Constituents of a California Straight-run Gasoline Distillate.

Page 2451. Column 2, first table, first column head should read "*d*-Camphonanic."—H. L. LOCHTE.

Louis F. Fieser. Synthesis of 2-Methyl-3-phytyl-1,4-naphthoquinone.

Page 2560. Column 1. "The last figure of the caption to Fig. 1 should read 3.42 in place of 2.42."—L. F. FIESER.

Daniel B. Clapp. The Reaction of a Thiophene Derivative with Maleic Anhydride.

Page 2735. In line 18, column 2, read "stilbene" for "maleic anhydride."—DANIEL B. CLAPP.

Roger Adams and E. F. Rogers. The Structure of Monocrotaline, the Alkaloid in *Crotalaria Spectabilis* and *Crotalaria Retusa*. I.

Page 2817. To Table II add:

	Alkaloid	Alkanol- amines	Acid
<i>Senecio</i>			
<i>integerrimus</i> ¹³	Integerrimine, C ₁₈ H ₂₃ O ₄ N Senecionine	Retronecine	Integerrinecic acid, C ₁₀ H ₁₄ O ₄
<i>spartioides</i> ¹³	Seneciphylline Spartioidine, C ₁₈ H ₂₃ O ₄ N		
<i>longilobus</i> ¹³	Longilobine, C ₁₈ H ₂₃ O ₄ N	Retronecine	Longinecic acid, C ₁₀ H ₁₄ O ₄
<i>ridellii</i>	Ridelliline, C ₁₈ H ₂₃ O ₄ N		
<i>Erechtites</i>			
<i>hieracifolia</i> ¹⁴	Hieracifoline, C ₁₈ H ₂₃ O ₄ N	Retronecine	Hieracinecic acid C ₁₀ H ₁₄ O ₄

—ROGER ADAMS.